



Chemical Methodologies

journal homepage: <http://chemmethod.com>



Original Research article

Silica Boron Sulfonic Acid as a New and Efficient Catalyst for the Green Synthesis of Quinoxaline Derivatives at Room Temperature

Sami Sajjadifar

Department of Chemistry, Payame Noor University, PO BOX 19395-4697 Tehran, Iran

ARTICLE INFORMATION

Received: 10 June 2017

Received in revised: 25 July 2017

Accepted: 10 August 2017

Available online: 24 August 2017

DOI:

10.22631/chemm.2017.88920.1000

KEYWORDS

Boron sulfonic acid (BSA)
quinoxaline synthesis
1,2-diamine, α -diketone
green chemistry

ABSTRACT

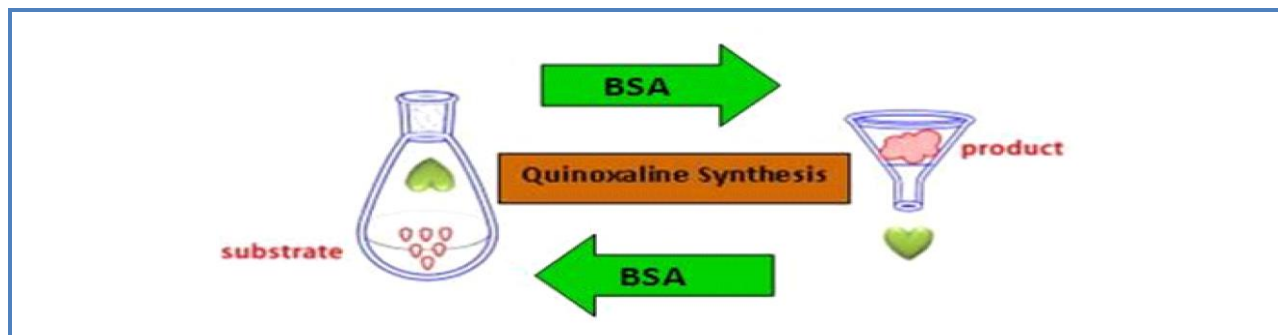
A simple, highly efficient and green procedure for the condensation of aryl and alkyl 1,2-diamines with α -diketones in the presence of catalytic amount of silica boron sulfonic acid (SBSA) at room temperature is described. By Using this method, quinoxaline derivatives as biologically interesting compounds are produced in high to excellent yields and short reaction times.

*Corresponding author, email: ss_sajjadifar@yahoo.com(S. Sajjadifar).

Department of Chemistry, Payame Noor University, PO BOX 19395-4697 Tehran, Iran

Tel.: +9884 32228316; Fax: +988432226101.

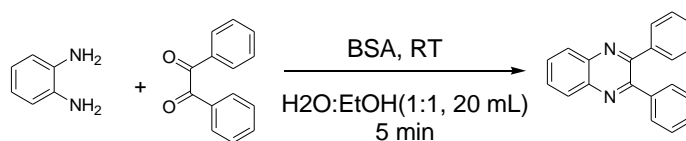
Graphical Abstract



Introduction

Quinoxaline derivatives have become increasingly important in the past few years because they have been proven to be extremely useful intermediates for the preparation of new biological materials [1-2]. Quinoxaline ring is a part of a number of antibiotics such as echinomycin, levomycin, and actinomycin, which are known to inhibit the growth of Gram-positive bacteria and are also active against various transplantable tumors [3].

Therefore, a number of synthetic strategies have been developed for the preparation of quinoxaline derivatives [4]. One of the most common methods is the condensation of 1,2-diamines with 1,2-dicarbonyl compounds in refluxing ethanol or acetic acid for 2–12h giving 34–85% yields [5]. Later, many improved methods have been reported for the synthesis of quinoxaline derivatives in the presence of various catalysts, such as Zn/L-proline [6], bismuth(III) triflate [7], metal hydrogen sulfates [8], molecular iodine [9], silica-bonded S-sulfonic acid [10], cerium(IV) ammonium nitrate [11], stannous chloride [12], zirconium tetrakis(dodecylsulfate) [13], amidosulfonic acid [14], montmorillonite K-10 [15], polyanilinesulfate salt [16], niobium pentachloride [17], Wells- Dawson heteropolyacid [18], ionic liquid [19], citric acid [20], ZrO_2/MxO_y ($M = Al, Ga, In, \text{ and } La$) mixed metal oxides supported on MCM-41 mesoporous molecular sieves [21]. The condensation has also been accomplished under catalyst-free conditions, but this needs microwave heating [22] in industry. An efficient catalyst-free protocol for the synthesis of quinoxaline derivatives was reported under ultrasound irradiation [23]. In recent years, significant articles have appeared reporting solid-state reactions by grinding [24-25]. Herein, we report a green, efficient, and convenient procedure for the synthesis of quinoxaline derivatives by the condensation of 1,2-diketones and 1,2-diamines using BSA as efficient and new catalyst at room temperature (Scheme 1).



Scheme 1. Quinoxaline synthesis by using BSA

In continuation of our studies on the application of BSA in organic synthesis [26-31], we describe our successful results that led to an extremely convenient method for quinoxalines synthesis in the presence of $\text{B}(\text{HSO}_4)_3$ as a solid acid catalyst in green condition and high isolated yields.

Materials and Methods

General

All reagents were purchased from Merck Fine Chemicals and were used without further purification. IR spectra of the compounds were obtained on a Shimadzu IR-435 spectrometer using a KBr disk. The ^1H nuclear magnetic resonance (^1H NMR) spectra were recorded on a Bruker AQS 300 Avance instrument at 300 MHz in CDCl_3 or dimethyl sulfoxide (DMSO-d_6) using tetramethylsilane as an internal standard. The progress of reaction was followed by a thin-layer chromatography (TLC) using silica gel SILG/UV 254 and 365 plates. New compounds and were characterized by IR, ^1H NMR, ^{13}C NMR, Mass spectroscopic data and their melting points.

General procedure for quinoxaline synthesis

1,2-diamine (1 mmol) was added to a mixture of α -diketone (1 mmol), boron sulfonic acid/ SiO_2 (0.03 g of mix., 0.03 mmol BSA) and $\text{EtOH}/\text{H}_2\text{O}$ [1/1 (v/v), 20 mL] in a 50 mL round-bottomed flask. The resulting mixture was stirred at room temperature for the times reported in Table 2. Then, H_2O (20 mL) was added to the reaction mixture and was allowed to stand at room temperature for 1 h. During this time, crystals of the pure product formed and were collected by filtration and dried under air.

Preparation of boron sulfonic acid (BSA) [26-39]

A 50 mL suction flask was equipped with a constant pressure dropping funnel. The gas outlet was connected to a vacuum system through an adsorbing solution (water) and an alkali trap. Boric acid (1.55 g, 25 mmol) was charged in the flask and chlorosulfonic acid (8.74 g, ca. 5 mL, 75 mmol in 5 mL CH_2Cl_2) was added dropwise over a period of 1 h at room temperature under $\text{N}_2(\text{g})$. HCl evolved immediately. After completion of the addition, the mixture was shaken for 85 min, while the residual HCl was eliminated by suction. Then the mixture was washed with diethyl ether to remove the unreacted chlorosulfonic acid. Then, we added 14.4 g silica gel and stirred them all. Finally, dried

and a grayish solid material was obtained (21.6 g, 95.66%). ¹H NMR of BSA in Acetone-D₆ show $\delta=12.218$. After isolation of the product, the filtrate was extracted with CHCl₃ (2×30 mL) and collect as organic layer. The aqueous layer (including oxalic acid) was separated, and its solvent was evaporated to obtain pure boron sulfonic acid. The recycled catalyst was used for the next run under identical reaction conditions.

Results and Discussion

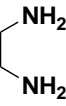
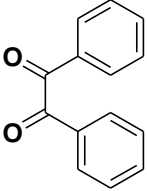
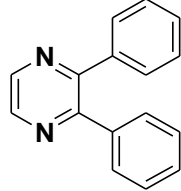
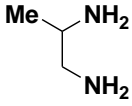
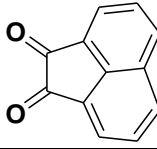
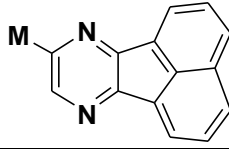
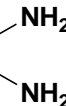
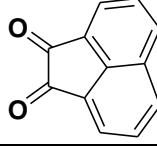
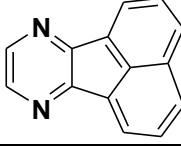
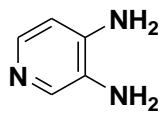
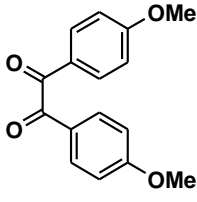
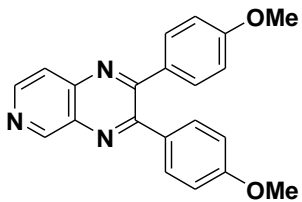
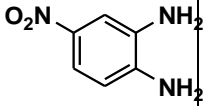
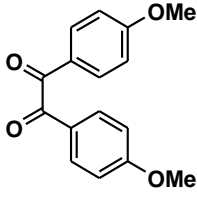
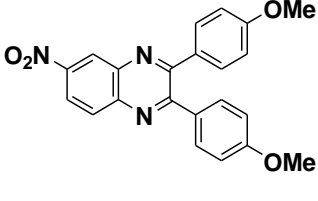
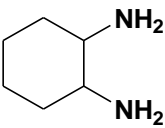
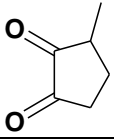
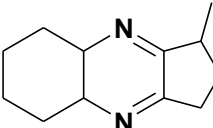
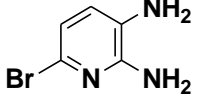

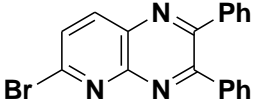
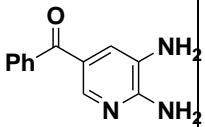
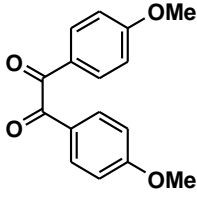
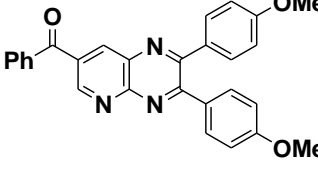
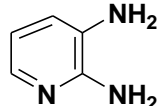
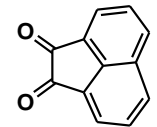
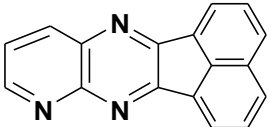
To optimize and select the best solvent for the reaction, the synthesis of quinoxaline **1a** was examined as a model in different solvents (Table 1). Higher yields and shorter reaction times were obtained when the reaction was carried out in EtOH:H₂O (1:1). Thus, EtOH:H₂O (1:1) was used as a reaction media for all reactions. Water is a desirable solvent for chemical reactions for a host of reasons such as cost, safety and environmental concerns. When 2,3- and 3,4-diaminopyridine derivatives reacted with 1,2-diketones, the yield decreased and reaction condition was harder (Table 2, Entries 1 and 7). In this research, *o*-phenylenediamines and 1,2-dicarbonyl compounds with electron-donating or electron-withdrawing groups were used. As indicated in the Table 2 both electron rich and electron deficient 1,2-dicarbonyl compounds worked pretty well, mostly leading to high yields of products.

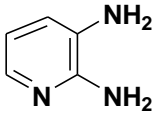
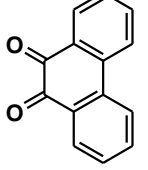
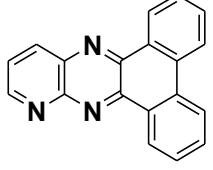
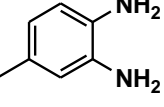
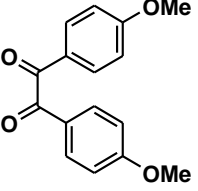
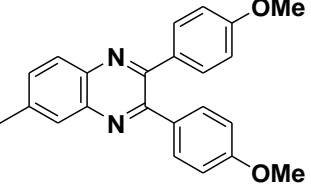
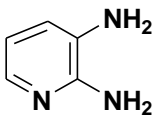
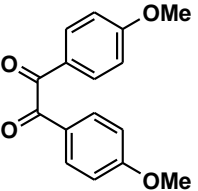
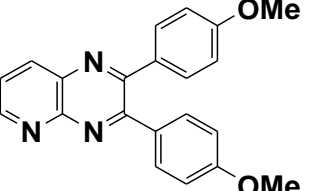
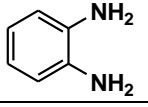
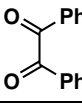
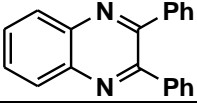
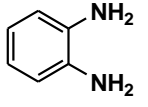
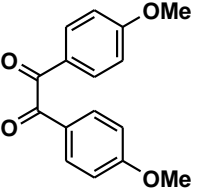
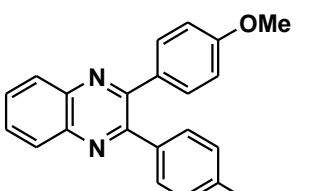
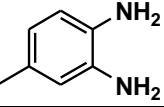
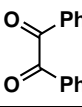
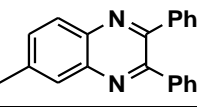
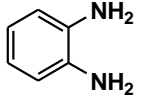
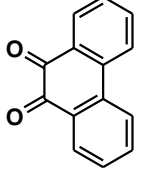
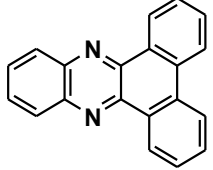
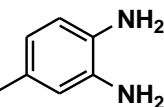
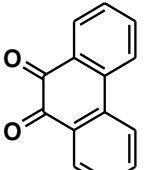
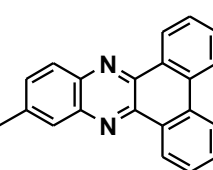
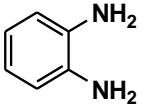
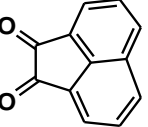
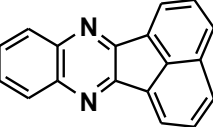
Table 1. The condensation of 1,2-diamine **1a** (1 mmol) with benzyl **2b** (1 mmol) in the presence of different ratios of BSA (0.03 mmol, 3 mol%) at room temperature

Entry	solvent	% BSA	Time (min)	Yield (%)
1	EtOH	5	30	78
2	EtOH	10	30	78
3	CH ₃ CN	5	60	72
4	CH ₃ COOEt	5	45	80
5	EtOH:H ₂ O (1:1)	15	10	81
6	EtOH:H ₂ O (1:1)	10	10	85
7	EtOH:H ₂ O (1:1)	5	10	85
8	EtOH:H ₂ O (1:1)	3	5	98
9	EtOH:H ₂ O (1:1)	2	25	87
10	EtOH:H ₂ O (1:2)	10	20	85
11	EtOH:H ₂ O (1:2)	5	20	85
12	EtOH:H ₂ O (1:2)	3	30	87
13	EtOH:H ₂ O (2:1)	3	25	84
14	H ₂ O	10	30	78
15	H ₂ O	5	30	78
16	CH ₂ Cl ₂	5	80	45
17	CHCl ₃	5	80	50

Table 2. Quinoxaline synthesis by using BSA (3 mol %) at room temperature

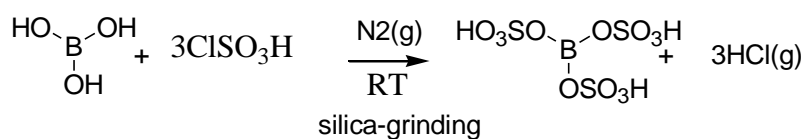
Entry	Diamine (DA)	Diketone (DK)	Product (Q)	Time (min)	Yield %	M.p°C
1				12 h	96	216-218
2				5	94	124-126
3				10	95	184-186
4				40	96	245-247
5				10	90	Liquid
6				5	98	58-61
7				24h	89	212-215
8				5	98	121-123

9				10	97	161-164
10				5	98	68-70
11				5	98	114-116
12				7h	93	120-122
13				3h	80	188-189
14				45	87	Oil
15				4.45h	90	143-145
16				41	86	145-147
17				5.5h	90	225-227

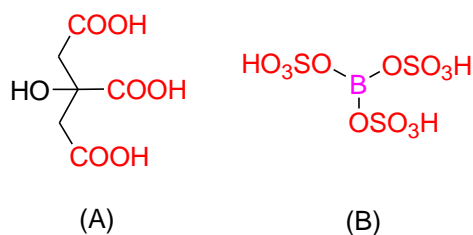
18				20	93	259-260
19				5	90	109-111
20				2.45h	90	109-111
21				3	98	125-127
22				30	89	134-136
23				15	97	113-111
24				3	96	224-226
25				5	97	219-221
26				10	99	241-242

27				10	99	231-233
28				12h	93	185-187

Boron sulfonic acid was easily prepared by adding chlorosulfonic acid to boric acid under N₂ atmosphere at room temperature. This reaction was easy and clean, because HCl gas was evolved from the reaction vessel immediately (Scheme 2) [32].

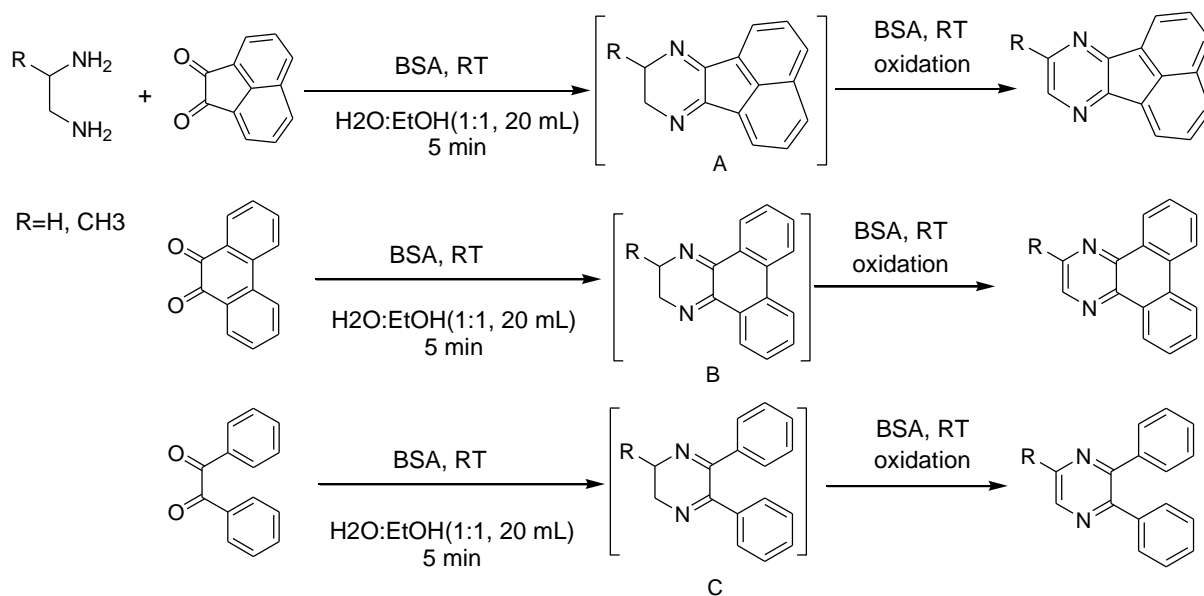


Scheme 2. Silica supported BSA production under N₂ atmosphere at room temperature



Scheme 3 A: Citric acid Structure and B: BSA Structure

Oxidation of A, B and C were accrued in presence of BSA catalyst (Scheme 3-B) but when citric acid (Scheme 3-A) was used as a weaker organo catalyst in similar condition, only A, B and C were obtained (Scheme 4). We think that the BSA acts as an acidic catalyst and oxidant agent (-SO₃H group). Similarly, when BSA was used as a catalyst for the synthesis of benzimidazole derivatives, oxidation of 2,3-dihydro-2-phenyl-1H-benzo[d]imidazole was occurred to corresponding product 2-phenyl-1H-benzo[d]imidazole.



Scheme 4. Oxidation in presence of BSA

In order to further validate our work, the current protocol was compared to the data in the literature based on the, temperature, reaction time, percentage yield, and loading catalyst.

The efficiency of various catalysts in synthesis of quinoxaline derivatives has been compared in Table 3. BSA as a catalyst afforded good results in comparison to the other catalysts.

Table 3. Comparison of catalytic ability of catalysts

Catalyst	Time	Catalyst loading	Yield ^c (%)	[Ref.]
Citric acid	1 min	10 mol%	94	[20]
SBSSA ^b	5 min	3.4 mol%	96	[10]
DMSO, I ₂	35 min	10 mol%)	95	[9]
Bentonite Clay K-10	20 min	2.5 g	95	[40]
BSA	3 min	3 mol%	98	This work

^aModel reaction is entry 21.

^bSilica Bonded S-Sulfonic Acid

^cIsolated yield.

To sum up our discussion, we have developed the use of BSA supported by silica gel as an inexpensive, easy to handle, non-corrosive and environmentally benign catalyst for the synthesis of quinoxalines from aromatic *o*-diamines and 1,2-dicarbonyl compounds. The advantages of the present procedure are simplicity of operation, short reaction times as compared to the other procedures for the preparation of quinoxaline derivatives, and the high yields of products. Moreover, in this reaction the catalyst can be recovered by filtration.

Acknowledgements

The authors gratefully acknowledge partial support of this work by Payame Noor University (PNU) of Iran.

References

1. Sakata G., and Makino K., Kuraswa Y. *Heterocycles*, 1988, **27**:2481
2. Kim Y.B., Kim Y.H., Park J.Y., Kim S.K. *Bioorg. Med. Chem. Lett.*, 2004, **14**:541
3. Dell A., William D.H., Morris H.R., Smith G.A., Feeney J. *J. Am. Chem. Soc.*, 1975, **97**:2497
4. Brown D.J. *In The Chemistry of Heterocyclic Compounds*; AQ2 Taylor, E. C., Wipf, P., Eds.; Wiley: New Jersey, 2004
5. Porter AEA. *In Comprehensive Heterocyclic Chemistry*; Katritzky AR, Rees CW, Eds.; Pergamon: Oxford, 1984, p 157
6. Heravi M.M., Tehrani M.H., Bakhtiari K., Oskooie H.A. *Catal. Commun.*, 2007, **8**:1341
7. Yadav J.S., Reddy B.V.S., Premalatha K., Shankar K.S. *Synthesis*, 2008, **23**:3787
8. Niknam K., Zolfigol M.A., Tavakoli Z., Heydari Z. *J. Chin. Chem. Soc.*, 2008, **55**:1373
9. Bhosale R.S., Sarda S.R., Ardhapure S.S., Jadhav W.N., Bhusare S.R., Pawar R.P. *Tetrahedron Lett.*, 2005, **46**:7183
10. Niknam K., Saberi D., Mohagheghnejad M. *Molecules*, 2009, **14**:1915
11. More S.V., Sastry M.N.V., Yao C.F. *Green. Chem.*, 2006, **8**:91
12. Shi D.Q., Dou G.L., Ni S.N., Shi J.W., Li X.Y. *J. Heterocycl. Chem.*, 2008, **45**:1797
13. Hasaninejad A., Zare A., Zolfigol M.A., Shekouhy M. *Synth. Commun.*, 2009, **39**:569
14. Li Z.J., Li W.S., Sun Y.J., Huang H., Ouyang P.K. *J. Heterocycl. Chem.*, 2008, **45**:285
15. Huang T.K., Wang R., Shi L., Lu X.X. *Catal. Commun.*, 2008, **9**:1143
16. Srinivas C., Kumar C., Rao V.J., Palaniappan S. *J. Mol. Catal A: Chem.*, 2007, **265**:227
17. Hou J.T., Liu Y.H., Zhang Z.H. *J. Heterocycl. Chem.*, 2010, **47**:703
18. Heravi M.M., Bakhtiari K., Bamoharram F.F., Tehrani M.H. *Monatsh. Chem.*, 2007, **138**:465
19. Potewar T.M., Ingale S.A., Srinivasan K.V. *Synth. Commun.*, 2008, **38**:3601
20. Mahesh R., Dhar A.K., Tara Sasank T.V.N.V., Thirunavukkarasu S., Devadoss T. *Chin. Chem. Lett.*, 2011, **22**:389
21. Ajaikumar S, Pandurangan A. *Appl. Catal. A: Gen.*, 2009, **357**:184
22. Zhao Z., Wisnoski D.D., Wolkenberg S.E., Leister W.H., Wang Y., Lindsley C.W. *Tetrahedron Lett.*, 2004, **45**:4873
23. Guo W.X., Jin H.L., Chen J.X., Chen F., Ding J.C., Wu H.Y. *J. Braz. Chem. Soc.*, 2009, **8**:1647

24. Tanaka K, Kishigami S, Toda, F. *J. Org. Chem.*,1991, **56**:4333
25. Kumar S., Sharma P., Kapoor K.K., Hundal M.S. *Tetrahedron*, 2008, **64**:536
26. Sajjadifar S. *Int. J. Chem. Tech. Res.*, 2013, **5**:385
27. Sajjadifar S., Khosravani E., Shiri S. *Int. J. Chem. Tech. Res.*,2013, **5**:1969
28. Sajjadifar S., Rezayati S. *Int. J. Chem.Tech. Res.*, 2013, **5**:1964
29. Sajjadifar S., Mirshokraie S.A., Javaherneshan N., Louie O. *Am. J. Org. Chem.*, 2012, **2**:1
30. Sajjadifar S. *Am. J. Org. Chem.*, 2012, **2**:116
31. Sajjadifar S., Louie O. *J. Chem.*, 2013, **2003**:No. 674946 (2013).
(<http://www.hindawi.com/journals/chem/2013/674946/>).
32. Kiasat A.R., Fallah-Mehrjardi M. *J. Braz. Chem. Soc.*, 2008, **19**:1595
33. Khazaeia A., Zolfigol M.A., Mokhlesi M., DerakhshanPanah F., Sajjadifar S. *Helvetica Chimica Acta*,2012, **95**:106
34. Madje B.R., Ubale M.B., Bharad J.V., Shingare M.S. *Bull. Cataly. Soci. Ind.*,2011, **9**:19
35. Mosslemin M.H., EshaghiMovahed A. *E-Journal of Chemistry*,2012, **9**:301
36. Parhami A., Khalafi-Nezhad A., Mowlazadeh Haghighi S., Bargebid R., Zare A., Moosavi-Zare A.R., Nikrooz M. *ARKIVOC*, 2012, **ix**:111
37. Sedighinia E., ZahedSargoli M. *Asian J. Chem.*, 2012, **23**:1456
38. Zolfigol M.A., Vahedi H., Massoudi A.H., Sajjadifar S., Loui O., Javaherneshan, N. *Clinical Biochemistry*, 2011, **44**:S219
39. Zolfigol M.A., Khazaei A., Vahedi H., Mokhlesi M., Sajadifar S., Pirveysian M. *Sulfur, and Silicon and the Related Elements*,2012, **187**:295
40. Hasaninejad A., Zare A., Shoekouhy M., Moosavi-Zare A.R. *E-Journal of Chemistry*, 2009, **6**:S247

How to cite this manuscript: Sami Sajjadifar. Silica Boron Sulfonic Acid as a New and Efficient Catalyst for the Green Synthesis of Quinoxaline Derivatives at Room Temperature. *Chemical Methodologies* 1(1), 2011, 1-11. DOI: [10.22631/chemm.2017.88920.1000](https://doi.org/10.22631/chemm.2017.88920.1000).